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SARKISYAN, Sergey Galustovich; MIKHAYLOVA, Nelli Aleksandrovna, Prinimali uchastiye: NIKITINA, R.G., nauchnyy sotr.; TROFIMU, I.A., nauchnyy sotr.; FEDOROV, S.F., otv. red.; STOLYAROV, A.G., red. izd-va; VOLKOVA, V.G., tekhn. red.

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(UKRAINE--TOMATOES--DISEASES AND PESTS) (MOSAIC DISEASES)
(PLANTS, EFFECT OF TEMPERATURE ON)

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Prinimali uchastiye: DAVYDOVSKAYA, N.G.; NIKITINA, R.Ya.; PILLER, Ya.Ya.

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NIKITINA, R.Ye.

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Zool. zhur. 44 no.2:294-296 '65. (MIRA 18:5)

1. Uzbekskiy nauchno-issledovatel'skiy veterinarnyy institut, Samarkandskaya oblast', selo Taylyak.

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Time of preservation of avian spirochetes in the organism of the hungry tick Argas persicus (Oken, 1818). Zool. zhur. 43 no.11:1723-1724 '64. (MIRA 18:11)

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MIKHLIN, E.D.; MEL'NIKOVA, G.K.; ZAYTSEVA, V.D.; NIKITINA, S.A.; GRITSMAN, Yu.Ya.; GORBOVITSKIY, Ye.B.; KRYUCHKOVA, G.S.; KONDRAT'YEVA, N.I.

Effect of rubber on drugs and the body. Report No.1: Present-day views on the subject. Med.prom. 12 no.2:35-41 F '58. (MIRA 11:3)

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(RUBBER--PHYSIOLOGICAL EFFECT) (DRUG INDUSTRY)

MIKHLIN, E.D., MEL'NIKOVA, G.K., ZAYTSEVA, V.D., NIKITINA, S.A., GRITSMAN,
Yu.Ya., GORBOVITSKIY, Ye.B., KRYUCHKOVA, G.S., KONDRAT'YEVA, N.I.

Effect of vulcanized rubber on drugs and the body. Report No.2.
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1. Nauchno-issledovatel'skiy institut reziny i Nauchno-issledovatel'skiy institut eksperimental'noy khirurgicheskoy apparatury i instrumentov.
(RUBBER--PHYSIOLOGICAL EFFECT)

TAUBMAN, A.B., doktor khimicheskikh nauk; NIKITINA, S.A.

Physical and chemical investigation of wetting agents used in dust catching. Bor'ba s sil. 2:61-70 '55. (MIRA 9:5)

1. Institut fizicheskoy khimii Akademii nauk SSSR.
(DUST--PREVENTION) (WETTING AGENTS)

NIKITINA, S.A.

2

The role of physicochemical factors in the interaction of dust particles with drops of solutions of surfactants. A. B. Taubman and S. A. Nikitina. *Doklady Akad. Nauk S.S.S.R.* 110, ~~[illegible]~~.—The dust-collecting ability of water and surfactant solns. was evaluated in a special app. by detg. nephelometrically the turbidity of dust suspensions escaping after spraying the suspended dust particles with a liquid that settles freely in the dust-filled space. Quartz dust, up to 20 μ, and coal and talcum dust, up to 5 μ in size, were settled with atomized water droplets contg. polyethylene glycols of alkylated phenols, Aerosol OT, tech. grade of sulfonol, isominol, 0.01% $CaCl_2$ soln., and alkylarylsulfonates. The results appear to indicate that the imperfect contact of the dust particles with water is not due to the air flowing around the water droplets but to the difficulties of wetting the solid surfaces and to the dynamic conditions during the brief contacts of dust with the droplets. The adsorption layer of the wetting agents has a hydrophilizing effect upon the solid particles and facilitates their contact with the liquid; this assures a better dust pptn. W. M. Sternberg

Inst. Phys. Chem AS USSR.

**NIKITINA, S.A.**

1-RG
4E2C

The mechanism underlying the capture of dust particles by water solutions of surface-active substances. A. B. Taubman and S. A. Nikitina. *Doklady Akad. Nauk S.S.S.R.* 110, 816-19(1956).—A study was made in connection with industrial and mining dust elimination. The capturing of dust particles by water soln. of various wetting agents is explained by a fast formation of the pure wetting agent over the surface of the solid particle, thus forming a tri-phase contact between the solid, the wetting agent, and the water.
N. Vasileff

3

Inst. Phys. Chem., AS USSR

NIKITINA, S A.

20-1-30/44

AUTHORS: Nikitina, S.A., Taubman, A.B.

TITLE: The Rôle of Adsorption Kinetics in the Effects of Wetting of Dust Particles by Drops of Solutions of Surface Active Substances (O roli kinetiki adsorptsii v yavleniyakh smachivaniya chastits pyli kaplyami rastvorov poverkhnostno-aktivnykh veshchestv)

PERIODICAL: Doklady AN SSSR, 1957, Vol. 116, Nr 1, pp. 113 - 116 (USSR)

ABSTRACT: When investigating the dust adsorbing effect of the aqueous solutions of surface active substances of semicolloidal type (e.g. the wetting substances ДБ and ОТ-10) the authors found considerable differences in their dust-adsorbing capacity, although these substances were of similar chemical composition and molecular structure. These differences were found also in such solutions in which a practically equal extremely reduced static surface tension (up to 30 - 35 erg/cm$^2$) exists. The reduction of the static surface tension of the solution of surface active substances does not characterize the true properties in the case of the real conditions of dust adsorption. The dynamic value $\sigma_d$ of surface tension may assume values that range nearly from the surface tension of pure water ( $\sigma_0$) up

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20-1-30/44

The Rôle of Adsorption Kinetics in the Effects of Wetting of Dust Particles by Drops of Solutions of Surface Active Substances

| | |
|---|---|
| ASSOCIATION: | Department for Dispersive Systems of the Institute for Physical Chemistry AN USSR<br>(Otdel dispersnykh sistem Instituta fizicheskoy khimii Akademii nauk SSSR) |
| PRESENTED: | April 13, 1957, by P.A. Rebinder, Academician |
| SUBMITTED: | April 4, 1957 |
| AVAILABLE: | Library of Congress |

Card 3/3

NIKITINA, S. A., Cand Chem Sci -- (diss) "Physico-chemical foundations of the dust-eliminating action of wetting agents." Mos, 1958. 16 pp. (Acad Sci USSR, Inst of Phys Chemistry), 100 copies. (KL, 9-58, 113)

**NIKITINA, S.A.**; TAUBMAN, A.B.

Apparatus for gauging the dust recovery properties of aqueous solutions of wetting agents. Gig. i san. 23 no.12:71-74 D '58. (MIRA 12:1)

1. Iz Instituta fizicheskoy khimii AN SSSR.
(DUST
dust recovery properties of aqueous solutions, appar. for evaluation (Rus))

NIKITINA, S.A., kand.khim.nauk; TAUBMAN, A.B., doktor khim.nauk; ZAKIYEVA, S.Kh., kand.khim.nauk

Physical and chemical principles of the dust-collecting action of surface active agents. Bor'ba s sil. 3:29-39 '59.
(MIRA 12:9)

(SURFACE ACTIVE AGENTS) (DUST--REMOVAL)

5(2), 21(7)

SOV/78-4-4-21/44

AUTHORS: Vdovenko, V. M., Lipovskiy, A. A., Nikitina, S. A.

TITLE: On the Formation of Chloride Complex Uranyl Compounds in Acetone (Ob obrazovanii khloridnykh kompleksnykh soyedineniy uranila v atsetone)

PERIODICAL: Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 4, pp 862-869 (USSR)

ABSTRACT: The formation of complex compounds of the uranyl ion in acetone with hydrochloric pyridine and hydroxylamine as $Cl^-$ donors was investigated by the spectrophotometric method. The authors plotted the absorption spectra of the solutions and the dependence of the molar extinction coefficient on the concentration of the complex-forming substance for various wavelengths (Fig 1). Uranyl perchlorate hexahydrate was applied as an initial compound. With a ratio of the components $UO_2(ClO_4)_2 : C_5H_5N.HCl = 1$ the complex $UO_2Cl_2$ is formed. With increasing concentration of hydrochloric pyridine or, more precisely, hydroxylamine also the complex formation is intensified. The complex $UO_2Cl_3^-$ is herein formed with hydrochloric

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TAUBMAN, A. B. and NIKITINA, S. A.

"The Kinetic Factor in Adsorption and Wetting Phenomena."

report presented at the Section on Colloid Chemistry, VIII Mendeleyev Conference of
General and Applied Chemistry, Moscow, 16-23 March 1959.
(Koll. Zhur. v. 21, No. 4, pp. 509-511)

VDOVENKO, V.M., LIPOVSKIY, A.A. NIKITINA, S.A.

Extraction of uranium from HCl solutions by means of tributyl phosphate. Zhur neorg. khim. 5 no.4:935-940 Ap '60.
(MIRA 13:7)

(Uranium) (Butyl phosphate)

S/020/60/135/005/035/043
B004/B075

AUTHORS: Taubman, A. B. and Nikitina, S. A.

TITLE: Micell Formation in Solutions of Surface-active Substances

PERIODICAL: Doklady Akademii nauk SSSR, 1960, Vol. 135, No. 5, pp. 1179-1182

TEXT: The authors discuss the problem 1) of the real existence of a critical concentration range of surface-active substances (SAS), in which micell formation takes place, and 2) of interpreting the critical points $C_{\delta r}$. The change of the solution properties of SAS was investigated. Methods and results are given in Table 1: Critical concentrations of micell formation:

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Micell Formation in Solutions of Surface-active Substances

S/020/60/135/005/035/043
B004/B075

| SAS preparation: | Methods of determining $C_{cr}$ | | | | | |
|---|---|---|---|---|---|---|
| | nephelometric | | titration with dye solution | | determination of surface tension | |
| | $(C_{cr})_1$ | $(C_{cr})_2$ | $(C_{cr})_1$ | $(C_{cr})_2$ | $(C_{cr})_1$ | $(C_{cr})_2$ |
| 1 | 0.0025 | 0.0049 | 0.0026 | 0.0050 | 0.0027 | 0.0052 |
| 2 | - | - | 0.014 | 0.031 | - | - |
| 3 | - | - | 0.00022 | 0.00045 | - | - |
| 4 | 0.0020 | 0.013 | 0.0024 | 0.012 | 0.0030 | 0.010 |
| 5 | 0.025 | 0.095 | 0.023 | 0.075 | 0.026 | 0.085 |

The authors examined: Preparation 1: Sodium salt of diethyl dihexyl ether of sulfosuccinic acid (aerosol OT); preparation 2: Sodium salt of dihexyl ether of sulfosuccinic acid (aerosol MA); preparation 3: Sodium salt of didecyl ether of sulfosuccinic acid; preparation 4: Dodecylamino acetate (amak 1120), and 5: preparation ДБ(DB), a polyoxy ethylene ester of ditertiary butyl phenol. The nephelometric determination was carried out by means of an ФЭК-Н-57 (FEK-N-57) nephelometer. In preparation 4, titration proceeded by eosin, in the others by pinacyanol chloride. For aerosol OT Fig. 1 shows the optical density D, the static surface tension $\sigma_s$, and

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Micell Formation in Solutions of Surface-active Substances S/020/60/135/005/035/043
B004/B075

the dynamic surface tension $\sigma_d$ as a function of concentration C. Two critical points were found: $(C_{cr})_1$ and $(C_{cr})_2$. The following interpretation is given: In the initial range of concentration $0 > C > (C_{cr})_1$ an adsorption layer forms, the SAS being in a molecular (ionic) state. With $C=(C_{cr})_1$ the maximum saturation of the adsorption layer and a minimum of the static surface tension ($\sigma_s = \sigma_{min}$) are obtained. With a further increase of C the excess molecules (ions) associate to micells, until $C = (C_{cr})_2$ is attained. From these results the conclusion is drawn that the micell formation as well as adsorption show the same tendency of long-chain saponaceous SAS, to decrease their free surface energy on the phase boundary. The authors thank Academician P. A. Rebinder for discussions. There are 1 figure, 1 table, and 15 references: 4 Soviet, 7 US, 3 British, and 1 German.

ASSOCIATION: Institut fizicheskoy khimii Akademii nauk SSSR (Institute of Physical Chemistry of the Academy of Sciences USSR)

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Micell Formation in Solutions of Surface-active Substances

S/020/60/135/005/035/043
B004/B075

PRESENTED:: July 5, 1960, by P. A. Rebinder, Academician

SUBMITTED: June 24, 1960

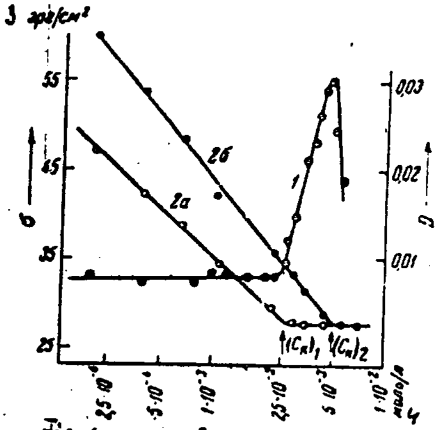

Fig. 1

Legend to Fig. 1: 1: D = f(C); 2a: $\sigma_s$ = f(C); 2b: $\sigma_d$ = f(C); 3: erg/cm$^2$; 4: mole/l.

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26602
S/186/61/003/004/001/007
E141/E164

21.4200

**AUTHORS:** Vdovenko, V.M., Lipovskiy, A.A., and Nikitina, S.A.

**TITLE:** On the mechanism of the extraction of tetravalent plutonium with primary alkylamine from $H_2SO_4$ solutions

**PERIODICAL:** Radiokhimiya, 1961, Vol.3, No.4, pp 396-402

**TEXT:** Extraction with high molecular weight aliphatic amines has recently found wide application. C.F. Coleman, K.B. Brown, J.G. Moore and K.A. Allen (Ref.1: Proceedings of the Second International Conference on the Peaceful Uses of Atomic Energy.) have described the extraction of $U^{IV}$, $Th^{IV}$, $Ce^{IV}$ and $Zr^{IV}$ from $H_2SO_4$ solutions, but they do not give any data on the mechanism of the reaction. The authors investigated this reaction mechanism. The authors used a mixture of primary aliphatic amines consisting of $C_7$ - $C_9$ atoms (d = 0.7846), the amount not exceeding 0.5%; chloroform was used as organic solvent. The $Pu^{IV}$ content in the aqueous and the organic phase was determined by measuring the α-activity of aliquot samples on a standard device type D. $Pu^{IV}$ was re-extracted from the organic phase into a 1.5N $HNO_3$ solution. Preliminary experiments showed that equilibrium was attained in not

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On the mechanism of the extraction ...

more than five minutes. The concentration of the amine in the starting solutions was determined by titrating with acid in an alcoholic solution whilst using bromophenolblue as indicator. After extraction the amine content in the aqueous and organic phase can be determined with alkalized solutions when the amine is converted into the base. The concentration of nitrate-ions in the organic phase was determined after re-extraction by titrating with indigo-red (Ref.6: J. Ungar, J. Appl. Chem., Vol.6 2, 245 (1956)). The sulphate-ion was determined gravimetrically as $BaSO_4$ and in some cases after $\beta$-irradiation with $S^{35}$ whilst using tagged $H_2SO_4$. When determining the distribution of the amine between the chloroform and the aqueous $H_2SO_4$ solution the sulphate was found mainly in the aqueous solution. The coefficients of distribution of plutonium ($\alpha$) increase with increasing concentration of the amine and of $H_2SO_4$ in the aqueous solution. This is possibly due to the formation of complex sulphate compounds of $Pu^{IV}$ containing the amine. The composition of the separated compound was determined by investigating the dependence of $\alpha_{Pu(IV)}$ on the concentration of the amine at constant concentration of $Pu^{IV}$ and $H_2SO_4$. It was found that two moles of amine nitrate

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On the mechanism of the extraction... S/186/61/003/004/001/007
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associate with one mole of the metal nitrate, forming the compound $(R_3NH)_2 Me (NO_3)_6$, where R = alkyl radical. The amine sulphate is not extracted by chloroform. The composition of the complex compounds extracted from $H_2SO_4$ solutions and from nitrate-sulphate solutions was confirmed by extraction and with the aid of absorption spectra. On the basis of concept on the anion-exchange character of extraction processes, carried out with amines, the separation of $Pu^{IV}$ from $H_2SO_4$ solutions can be envisaged as a sulphate-ion exchange on an anion complex of $Pu^{IV}$ which is formed in the aqueous phase:

$$Pu^{4+} + 3SO_4^{2-} \rightleftarrows Pu(SO_4)_3^{2-}$$

$$2(RNH_3)_2SO_4 + Pu(SO_4)_3^{2-} \rightleftarrows (RNH_3)_4Pu(SO_4)_4 + SO_4^{2-} \qquad (1)$$

The extraction can also be described by the formation of neutral plutonium sulphate which is formed from the amine sulphate of the extracted complex compound:

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On the mechanism of the extraction ... S/186/61/003/004/001/007
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$$Pu^{4+} + 2SO_4^{2-} \rightleftarrows Pu(SO_4)_2$$

$$2(RNH_3)_2SO_4 + Pu(SO_4)_2 \rightleftarrows (RNH_3)_4Pu(SO_4)_4 \qquad (2)$$

$Pu^{IV}$ is known to form complex sulphate compounds even when a considerable excess of nitrate-ions is present in the solution (Ref.5: A.A. Lipovskiy, N.B. Chernyavskaya, ZhNKh, Vol.4, 10, 2244 (1959)). Experiments were therefore carried out on the extraction of plutonium from a mixture of $HNO_3$ and $H_2SO_4$ solutions. At low concentrations of the amine the nitrate is hardly extracted by chloroform but with increasing concentration polymeric amine aggregates are formed which are extracted by the organic solvent. An increase in the concentration of $HNO_3$ in the aqueous solution at constant concentration of the amine leads to an increased content of the amine nitrate in the organic phase. Colloidal solutions of amine nitrate in chloroform show a different behaviour on extraction with respect to ions which are present in the aqueous solution and experiments were carried out to investigate the behaviour of the sulphate-ion during extraction from nitrate-sulphate solutions. When the concentration of $HNO_3$, and therefore
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E141/E164

also of the amine nitrate in the solution was small, the amine sulphate was not extracted into the chloroform. Deposits were formed at the phase boundary as well as during extraction from pure $H_2SO_4$ solutions. At concentrations of approximately 0.15N $HNO_3$ the deposits disappear and the concentration of the sulphate-ion in the organic phase increases sharply. A gradual substitution of the sulphate-ions and nitrate-ions takes place during the later stages. It is also suggested that the amine sulphate is separated from the nitrate-sulphate solutions because a finely dispersed phase is present in the chloroform, the dispersion being formed by the amine nitrate. Acknowledgments are expressed to L.N. Lazarev for his cooperation.

There are 5 figures, 2 tables and 7 references: 2 Soviet and 5 non-Soviet. The 4 English language references read as follows:

Ref.1: as in text above.
Ref.3: A.S. Wilson. Proceedings of the Second International Conference on the Peaceful Uses of Atomic Energy.
Ref.6: As in text above.
Ref.4: D.J. Carswell, J.J. Laurenset. J.Inorg.Nucl.Chem., V.11, 1, 69 (1959)

Card 5/5. SUBMITTED July 5, 1960.

NIKITINA, S.A.; KONSTANTINOVA, V.V.; ZAKIYEVA, S.Kh.; TAUBMAN, A.B.

Wetting capacity of surface-active substances and their rate of adsorption from aqueous solutions. Zhur. prikl. khim. 34 no.12:2658-2664 D '61. (MIRA 15:1)

(Surface-active agents) (Adsorption)

FILATOVA, M.A.; NIKITINA, S.A.; TAUBMAN, A.B.; REBINDER, P.A., akademik

Demulsifying effect of surface active agents and the structural and mechanical properties of their adsorption layers. Dokl. AN SSSR 140 no.4:874-876 O '61. (MIRA 14:9)

1. Institut fizicheskoy khimii AN SSSR.
(Surface active agents) (Emulsions)

S/186/62/004/006/001/009
E075/E433

AUTHORS: Vdovenko, V.M., Lipovskiy, A.A., Nikitina, S.A.

TITLE: On the mechanism of extraction of $U^{(VI)}$ with solutions of tridecylamine fluoride

PERIODICAL: Radiokhimiya, v.4, no.6, 1962, 625-632

TEXT: In order to elucidate the mechanism of the extraction process, some relationships in the extraction of $U^{(VI)}$ from solutions in HF and NaF with tridecylamine fluoride (TDA) were studied. Using a spectrographic method it was shown that in organic solutions complex compounds $TDAHUO_2F_3$ and $(TDAH)_2UO_2F_4$ are formed. The composition of the complex compounds was also determined from the results of the analysis of equilibrium organic solutions and construction of the dependence $\lg \alpha_U$ on $\lg C_{TDA \cdot HF}$ (where $\alpha_U$ - the coefficient of distribution of uranium). It was shown that complex compounds with the ratio $F/U > 4$ were not formed. The investigation of the extraction of uranyl fluoride from aqueous solutions with tridecylamine fluoride indicated that the extraction takes place due to the formation of complexes of $UO_2F_2$ with one or two molecules of TDA·HF. The
Card 1/2

On the mechanism ...

S/186/62/004/006/001/009
E075/E433

best removal of $U^{(VI)}$ with amine fluoride is obtained on extraction from solutions of $UO_2F_2$ not containing an excess of fluorine ions. There are 5 figures and 7 tables.

SUBMITTED: August 26, 1961

Card 2/2

S/186/63/005/001/011/013
E075/E436

AUTHORS: Vdovenko, V.M., Lipovskiy, A.A., Nikitina, S.A.

TITLE: Investigation of the solvation of uranyl chloride with the molecules of tri-n-butyl phosphate

PERIODICAL: Radiokhimiya, v.5, no.1, 1963, 139-141

TEXT: Spectroscopic methods were used to investigate the nature of $UO_2Cl_2$ added to a solution of tri-n-butylphosphate (TBP) in $CCl_4$. In the original solution (saturated solution of anhydrous $UO_2Cl_2$ in TBP dissolved in $CCl_4$) the solute had the composition of $UO_2Cl_2 \cdot 2TBP$. As the proportion of TBP increased (100% TBP), the UV spectra of the solutions changed and indicated that $UO_2Cl_2$ coordinates with 3 molecules of TBP. The equilibrium constant for the formation of $UO_2Cl_2 \cdot 3TBP$ is 3.1 ± 0.2. The interaction of $P = O$ groups with the U atom was shown in infrared spectra to be less strong in $UO_2Cl_2 \cdot 3TBP$ than in $UO_2Cl_2 \cdot 2TBP$. The complex with 4 molecules of TBP did not form, presumably due to steric hindrance. There are 3 figures and 1 table.

SUBMITTED: October 31, 1962

Card 1/1

VDOVENKO, V.M.; LIPOVSKIY, A.A.; NIKITINA, S.A.

Study of the solvation of $UO_2Cl_2$ with molecules of organophosphorus compounds by spectral methods. Radiokhimiia 5 no.5:585-591 '63.
(MIRA 17:3)

NIKITINA, S.A.; SPIRIDONOVA, V.A.

Colloid chemical properties of solutions of sulfonol NP and its analogs and their detergent action. Koll.zhur. 25 no.5:581-586 S-O '63.
(MIRA 16:10)

1. Institut fizicheskoy khimii AN SSSR, Moskva.

LEVITSKIY, L.M.; NIKITINA, S.A.; TAUBMAN, A.B.

Formation of micelles in solutions of surface-active substances and kinetic wetting. Dokl.AN SSSR 149 no.3:633-636 Mr '63. (MIRA 16:4)

1. Institut fizicheskoy khimii AN SSSR. Predstavleno akademikom P.A.Rebinderom.
(Micelles) (Surface-active agents)

S/020/63/149/004/022/025
B106/B186

AUTHORS: Nikitina, S. A., Taubman, A. B., Kuligina, N. V., Spiridonova, V. A.

TITLE: Structuration in interphase adsorption layers of solutions of surface-active substances and stability of emulsions and aqueous dispersions of polymers (latex)

PERIODICAL: Akademiya nauk SSSR. Doklady, v. 149, no. 4, 1963, 905 - 908

TEXT: The values of the shear stress $P_m$ of the adsorption layers in aqueous ОП-10 (OP-10) (alkyl phenol polyhydroxyethylene ether) solutions during the period of their formation under static conditions were measured along the water - xylene interface. Furthermore, the stability of concentrated emulsions of xylene in water and polymer dispersions (polystyrene latex stabilized with OP-10 and pluronic (hydroxy ethylene and hydroxy propylene block polymer)) was studied. The rate of increase in strength of the protective emulsifier layers increases rapidly with the concentration of OP-10 solutions. Even in 5% solutions, however, maximum strength ($P_m$ = 0.25 dyn/cm) is only reached after 25 hrs. When the same amount of OP-10 was previously distributed between the two phases, high strength was

Card 1/3

Structuration in interphase... S/020/63/149/004/022/025
B106/B186

reached in the first few minutes after the interface formed. To obtain small droplets (2 - 3μ) the emulsion has to be shaken vigorously for 15 - 20 min. In a 5% solution of pluronic L64 the layer solidifies immediately even if the surface-active substance is dissolved in the aqueous phase only. The drop size is 2 - 3μ also after short-time shaking. These facts confirm the relation between high strength of protective layers ($P_m \approx 0.1$ dyn/cm) and the formation of complex supermolecular structures in the form of multilayer phase layers on the interface of the two liquid phases. These structures can also be seen visually. The above phenomena are caused by the hydrodynamic effect of the spontaneous surface turbulence and convection, which causes an unidirectional transition of the hydrocarbon phase into the aqueous phase in the form of an ultramicroemulsion (C. Sternling, L. Scriven, Am. Inst. Chem. Eng., 6, 514 (1959)). The results obtained by using OP-10 and pluronic 164 as emulsifiers for polystyrene latex agree well with the above data. These emulsifiers as well as Nekal'-НБ (Nekal'-NB) and pluronic 164 make an almost complete polymerization of the monomer possible at comparatively high concentrations, but differ considerably in their stabilizing action. Maximum stability is reached when using OP-10 or pluronic as emulsifiers under conditions such

Card 2/3

S/020/63/149/004/022/025
Structuration in interphase... B106/B186

that solid supermolecular surface structures are able to form. High strength of the structurized stabilizing layers is the principal condition for high stabilization of latex and concentrated emulsions; it is achieved when the emulsifiers are sufficiently soluble in water as well as in the non-aqueous phase. There are 1 figure and 1 table. The most important English-language reference is: A. Kaminski, J. W. McBain, Proc. Roy. Soc., London, A198, 447 (1949).

ASSOCIATION: Institut fizicheskoy khimii Akademii nauk SSSR (Institute of Physical Chemistry of the Academy of Sciences USSR)

PRESENTED: September 26, 1962, by P. A. Rebinder, Academician

SUBMITTED: September 12, 1962

Card 3/3

VDOVENKO, V.M.; LIPOVSKIY, A.A.; NIKITINA, S.A.

Hydrogen bonding in alkyl ammonium salts. Part 2: Infrared spectra and structure of tridecyl ammonium chloride. Radiokhimiia 6 no. 1:56-62 '64. (MIRA 17:6)

VDOVENKO, V.M. (Leningrad); [illegible], A.A. (Leningrad); KUZINA, M.G. (Leningrad); DEM'YANOVA, T.A. (Leningrad); NIKITINA, S.A. (Leningrad)

Hydrogen bonds in alkyl ammonium salts. Ukr. fiz. zhur. 9 no.4:453-457 Ap '64. (MIRA 17:8)

L 41058-65 EPF(c)/EWP(j)/EWT(m) Pc-4/Pr-4 RM
ACCESSION NR: AP5007141 S/0303/65/000/001/0020/0023 24
21
B

AUTHOR: Nikitina, S.A.; Spiridonova, V.A.; Taubman, A.B.

TITLE: Stabilization of polyvinylacetate latexes by nonionogenic emulsifiers

SOURCE: Lakokrasochnyye materialy i ikh primeneniye, no. 1, 1965, 20-23

TOPIC TAGS: latex stabilization, polyvinylacetate, latex, latex emulsifier, nonionogenic emulsifier, surfactant polymer, ethylene oxide copolymer, propylene oxide copolymer, film strength, ethylenediamine copolymer, latex stability, polyvinyl film/proxanol, proxamine

ABSTRACT: American made copolymers of ethylene oxide and propylene oxide, American made ethylenediamine-based copolymers, and similar Soviet-made compounds (proxanols-2, 3, 4, 5, 6, and proxamines-1 through 6) were the nonionogenic surface-active substances used by the authors as emulsifiers in the synthesis of polyvinylacetate latexes. The 2-4 hr. polymerization at 58-60C, with the phase ratio between polymer and aqueous solution of emulsifier maintained at 1:1, was carried out in nitrogen-filled ampules in a constant temperature water bath, using potassium persulfate (0.08% of polymer weight as the reaction initiator. The low-temperature stability of the obtained latexes was investigated according to a method developed at VNIISK, i.e., repeated 1-hr freezing of the samples at -25C, sedimentation

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L 41058-65
ACCESSION NR: AP5007141 3

at room temperature, 1-hr. centrifugation at 4000 rpm and determining the dry residue in the separated coagulates. Tensile strengths of 34 to 117.8 kg/cm$^2$ with a relative elongation of 200 to 980% were obtained in polyvinyl films prepared with proxanol 3 and 6 and proxamine 2 and 3, indicating that highly stable latexes yielding polymer films 15 with adequate physical properties can be prepared using these emulsifiers. "The authors thank K. M. Mizuch and Ts. M. Gel'fer, candidates in the chemical sciences, who provided the emulsifier preparations." Orig. art. has: 5 tables and 2 figures.

ASSOCIATION: none

SUBMITTED: 00 ENCL: 00 SUB CODE: MT, OC

NO REF SOV: 011 OTHER: 006

Card 2/2

L 17376-66 EPF(n)-2/EWT(m)/EWP(t) IJP(c) WW/JD/JG

ACC NR: AP6004504 SOURCE CODE: UR/0186/65/007/005/0509/0516

AUTHOR: Vdovenko, V. M.; Lipovskiy, A. A.; Nikitina, S. A.; Yakovleva, N. Ye.

ORG: none

40
B

TITLE: Investigation of the extraction of $U^{IV}$ and $U^{VI}$ from hydrochloric acid solutions by means of tri-n-butylphosphate

SOURCE: Radiokhimiya, v. 7, no. 5, 1965, 509-516

TOPIC TAGS: uranium, organic phosphorus compound, solvent extraction, complex molecule

21,55

ABSTRACT: The uranium was extracted from the aqueous phase by forming the complex compounds which accumulated in the organic phase. The optical method (percent transmission of 400-700 millimicrons) was applied to measurement of the concentration of uranium-tri-n-butylphosphate complexes in the organic phase. The extractions were conducted using either 20% in $CCl_4$ or 100% TBP. In the extraction experiments 0.5-12.8 molac HCl solutions and 5-10.9 molar LiCl solutions were used. It was found that the composition of the complexes formed is a function of both the

UDC: 542.61:546.791.4$^{2}$791.6

2

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L 17376-66

ACC NR: AP6004504

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HCl concentration in the aqueous phase and the TBP concentration in the inert solvent. In the case of $U^{VI}$, the following complexes were found in the extracts: $UO_2Cl_2(TBP)_2$, $UO_2Cl_2(TBP)_3$, and a complex anion $[UO_2Cl_3(TBP)_n]^-$. In the case of $U^{IV}$, the organic phase contained $UCl_4(TBP)_2$, $UCl_4(TBP)_3$, and a complex anion $UCl_6^{2-}$. Under the conditions near saturation equilibrium, both the $U^{IV}$ and the $U^{VI}$ are combined with two molecules of TBP. In the case of an excess of TBT, the complex involves three molecules of TBP. In the case of higher HCl concentration in the starting aqueous solution, accompanied by an excess of TBP, the extract contains anionic complexes of $U^{IV}$ and $U^{VI}$. Orig. art. has: 2 figures, 2 tables, 6 formulas.

SUB CODE: 07/ SUBM DATE: 02Nov64/ ORIG REF: 013/ OTH REF: 006

Card 2/2 nst

L 17375-66 EWP(j)/EWT(m)/T RM

ACC NR: AP6004505 SOURCE CODE: UR/0186/65/007/005/0563/0572

AUTHOR: Lipovskiy, A. A.; Nikitina, S. A.; Yakovleva, N. Ye.

39

37

B

ORG: none

TITLE: Investigation of the $UCl_4$ solvation by molecules of neutral organophosphate compounds by means of spectroscopic methods

SOURCE: Radiokhimiya, v. 7, no. 5, 1965, 563-572

TOPIC TAGS: uranium compound, organic phosphorous compound, complex molecule, solvent action, intermolecular complex, IR spectrometer, absorption spectrum

ABSTRACT: Solvation of $UCl_4$ by tri-n-butylphosphate (TBP), diisoamine ester of methylphosphonic acid (DAMPA), and tributylphosphineoxide (TBPO) was investigated using optical methods. Absorption spectra were taken with an SF-2M spectrophotometer (400-1100 millimicrons) and IKS-14IR spectrometer (7-11 millimicrons). Saturated solutions of anhydrous $UCl_4$ in solvents containing TBP-, DAMPA-, and TBPO in $CCl_4$ and benzene were used. It was found that in the case of an excess of phosphoorganic ligands in inert solvents, the complexes of the general formula $UCl_4S_3$ are

UDC: 548.56 : 546.791.4'131

Z

Card 1/2

L 17375-66
ACC NR: AP6004505

2

formed (where S is TBP, DAMPA, or TBPO). The equilibrium contents of the reaction $UCl_4S_2 + S \rightleftarrows UCl_4S_3$ are 55±3, 117±12, and 86±4 for TBP, DAMPA, and TBPO, respectively. Changes in the absorption spectra of $U^{IV}$ which occur at high DAMPA- and TBPO concentrations are attributed to the coordinatively saturated compounds of the $UCl_4S_4$ type. The increase in ligands concentration was found to be reflected in weaker bonds between $U^{VI}$, and $U^{IV}$ atoms and phosphoorganic ligands 14 in $UCl_4S_3$ type complexes. It was found that the electron affinity of heavy metal salts can be measured in terms of displacement of the IR absorption spectra of vibration of P=O group of the coordination compounds containing equal number of neutral ligands. The similarity of the absorption spectra of $UCl_6^{2+}$ and $UCl_4S_2$ indicate that the ligands are arranged octahedrally around the $U^{4+}$ ion in the $UCl_4S_2$ complex. A low symmetry is assigned to the $UCl_4S_3$ compound. The absorption spectra are graphed. Orig. art. has: 5 figures, 3 tables.

SUB CODE: 07/ SUBM DATE: 02Nov64/ ORIG REF: 010/ OTH REF: 009

Card 2/2 nst

TAUBMAN, A.B.; NIKITINA, S.A.; PRIGORODOV, V.N.

Role of quasi-spontaneous emulsification in the process of stabilization of emulsions. Koll. zhur. 27 no.2:291-292 Mr-Ap '65. (MIRA 18:6)

1. Institut fizicheskoy khimii AN SSSR, Moskva.

[illegible], V.N.; NIKITINA, S.A.; TAUBMAN, A.B.

Stability of concentrated emulsions [illegible] emulsifiers. Koll. zhur. [illegible]

(MIRA [illegible])

1. Institut fizicheskoy khimii AN SSSR, Moskva. Submitted July 1[illegible], 1961.

ACC NR: AP7006039 SOURCE CODE: UR/0063/66/011/004/0387/0391

TAUBMAN, A. B., Doctor of Chemical Sciences, NIKITINA, S. A., Candidate of Chemical Sciences, and TOLSTAYA, S. N., Candidate of Chemical Sciences

ORG: none

"Surface-active Agents in the Physical Chemistry and Technology of Polymers"

Moscow, Zhurnal Vsesoyuznogo Khimicheskogo Obshchestva, Vol 11, No 4, 1966, pp 387-391

Abstract: With the exception of carbon black -- a highly hydrophobic and highly dispersed and therefore effective polymer filler, most other mineral fillers, including pigments, are hydrophilic and of low activity. The oriented adsorption of surface-active agents (SAA) on particle surfaces, lyophilizing it with respect to the polymeric medium (the polymer or its solution in an organic solvent), peptizes the filler, intensifies the molecular interaction on the particle-medium interface, and alters the strength properties arising in the filling of coagulational structures. This kind of modification with the purpose of facilitating the distribution of filler particles in the polymer (rubbers), improving the wear of pigments, and increasing the stability of paints is carried out with the aid of long-chain amines, fatty acids, their salts, and many other SAA. The high emulsifying capacity of long-chain SAA homologs -- the soaps of natural fatty acids and synthetic soap-like compounds -- affords their extensive utilization in the polymerization of latexes. For this purpose, anion-active and non-ionogenic soaps are chiefly used. Orig. art. has: 1 figure. [JPRS: [illegible]]

UDC: 661.185.1 + 541.6

Card 1/2 0927 0828

ACC NR: AP7006039

TOPIC TAGS: surface active agent, pigment

SUB CODE: 11 / SUBM DATE: none / ORIG REF: 051 / OTH REF: 020

Card 2/2

VDOVENKO, V.M.; LIPOVSKIY, A.A.; NIKITINA, S.A.; YAKOVLEVA, N.Ye.

Extraction of $U^{(IV)}$ and $U^{(VI)}$ from hydrochloric solutions by means of tri-n-butyl phosphate. Radiokhimiia 7 no.5:509-516 '65. (MIRA 18:10) [illegible], YAKOVLEV, N.Ya.

[illegible] method of investigation of the solvation between $HCl_4$ and molecules of neutral organophosphorus compounds. Radiokhimiia [illegible] no.5:563-572 '65. (MIRA 18:10)

LIPOVSKIY, A.A.; NIKITINA, S.A.

Infrared spectra and structure of trialkyl ammonium fluorides. Zhur. neorg. khim. 10 no.1:176-182 Ja '65.
(MIRA 18:11)

1. Submitted July 18, 1963.

ARTAMONOV, Ye.V.; NIKITINA, [illegible].B.; CHECHURINA, Ye.N.

Determination of the hysteresis curves of ferromagnetic materials.
Nov.nauch.-issl.rab.po.metr. VNIIM no.5:4-6 '64.
(MIRA 18:3)

NIKITINA, [illegible]; [illegible]

Determining the [illegible] steels. Nov.nau. [illegible] .rab.; .met. [illegible]
(MIRA 18:3)

SVERCHINSKIY, B.S.; MOLOKANOV, Yu.K.; NIKITINA, S.D.; PRIGOZHINA, L.D.

Determining the coefficients of relative volatility of component pairs in a multicomponent mixture. Zhur. fiz. khim. 39 no.9:2117-2119 S '65. (MIRA 18:10)

L 30404-66 [illegible] T1 IJP(c) RM/DS/WW/JW/JD/JG

ACC NR: AP60[illegible] SOURCE CODE: UR/0076/66/040/002/0480/0481

AUTHOR: Vas[illegible]

ORG: Moscow [illegible] M. V. Lomonosov (Moskovskiy gosudarstvennyy universitet)

TITLE: Therm[illegible] W – O system

SOURCE: Zhurn[illegible] v. 40, no. 2, 1966, 480-481

TOPIC TAGS: [illegible], tungstate, thermodynamic function, equilibrium constant, chemical reduction, [illegible], x ray analysis

ABSTRACT: The reduction of CaWO4 with hydrogen was investigated by using the circulation method of studying heterogeneous equilibria, and the products were analyzed by x-ray phase analysis. The reduction process at 950 – 1170C was found to take place in three stages. The first stage is characterized by a slight decrease of the equilibrium constant with a decrease in oxygen content, apparently within the confines of the region of homogeneity of $CaWO_4$; the boundaries of this region expand somewhat with rising temperature. The second stage is characterized by the fact that the $K_p$ values are independent of the composition up to $CaWO_2$ (total composition). The presence of the third stage is deduced from a lower value of the equilibrium constant for tungstates with a lower oxygen content than that of $CaWO_2$. X-ray phase analysis showed that during the first stage the reduction products contain only one phase, $CaWO_{4-x}$, with an undistorted $CaWO_4$ lattice; during the second stage, the reduction

UDC: 541.11

Card 1/2

L 30404-66

ACC NR: AP6008097

4

products consist of a mixture of the phases $CaWO_{4-x}$, W, and $Ca_3WO_6$. Values of the partial thermodynamic functions $\Delta G_T$, $\Delta H$, and $\Delta S$ (per mole of $O_2$) pertaining to the mixture of the phases $Ca_3WO_6$, W, and $CaWO_{4-x}$ were calculated. The authors are deeply grateful to L.M. Kovba for assistance in carrying out the x-ray phase analysis.

SUB CODE: 07 / SUBM DATE: 16Jul65 / ORIG REF: 002 / OTH REF: 001

Card 2/2 CC

MEREZHINSKIY, M.F.; NIKITINA, S.M.

Adaptation of the animal body under conditions of different consumption of fats with saturated and unsaturated fatty acids. Vop. pit. 21 no.3:37-40 My-Je '62. (MIRA 15:10)

1. Iz kafedry biokhimii (zav. - prof. M.F.Merezhinskiy) Minskogo meditsinkogo instituta.
(FAT METABOLISM) (CHOLESTEROL) (ASCORBIC ACID)

NIKITINA, S.S.

Mineralogy of basic effusive rocks of Marmarosh massif. S. S. Nikitina (Lvov Univ.). *Mineralog. Sbornik, L'vov. Geol. Obshchestvo* 4, 163-6(1950).—Upper Jurassic diabase rocks intruding and capping ancient cryst. rocks of the massif are composed principally of plagioclase and pyroxene. Accessory minerals are ilmenite, magnetite, and apatite, and secondary minerals are calcite, lotrite (pumpellyite), chlorite, and epidote. The presence of lotrite in relatively significant amounts indicates that the rocks were subjected to low-temp. metamorphism at shallow depth. G P

Marie Siegrist

SHAPIRO, V.M.; NIKITINA, S.S.; VOLKOVA, A.D.

Effect of intravenous introduction of sodium bicarbonate on the acid-base equilibrium and gases in the blood during direct cardiac massage. Pat. fiziol. i eksp. terap. 8 no.1:29-33 Ja-F '64. (MIRA 1[illegible]

1. Laboratoriya eksperimental'noy fiziologii po ozhivleniyu organizma (zav.- prof. V.A. Negovskiy) AMN SSSR, Moskva.

RABINOVICH, I.M.; KIBAL'CHICH, P.N.; FADEYEVA, I.I.; IL'INSKAYA, T.N.;
KUZOVKOV, A.D.; BEREZHINSKAYA, V.V.; TRUTNEVA, Ye.A.; NIKITINA, S.S.

Plants of the Stephania genus as a source of new medicinal preparations. Apt. delo 14 no.6:19-22 N-D '65. (MIRA 18:12)

1. Vsesoyuznyy nauchno-issledovatel'skiy institut lekarstvennykh i aromaticheskikh rasteniy, Moskva. Submitted June 15, 1965.

BEREZHINSKAYA, V.V.; NIKITINA, S.S.

Anti-inflammatory effect of tetrandrin and some aspects of its mechanism of action. Farm. i toks. 28 no.1:77-81 Ja-F '65.
(MIRA 18:12)

1. Laboratoriya narodnoy meditsiny (zav. - kand.med.nauk V.V.Berezhinskaya) Vsesoyuznogo nauchno-issledovatel'skogo instituta lekarstvennykh i aromaticheskikh rasteniy, Moskva. Submitted August 2, 1963.

KUZ'MIN, R.N.; NIKITINA, S.V.

Structure of compounds of rare earth metals with antimony and bismuth of an AB composition. Kristallografiia 8 no.3:453-454 My-Je '63. (MIRA 16:11)

1. Moskovskiy gosudarstvennyy universitet imeni Lomonosova.

KUL'SKIY, L.A.; NIKITINA, S.V.

Determination of silver in oligodynamic solutions.
Ukr.khim.zhur. 28 no.8:977-980 '62. (MIRA 15:11)

1. Institut obshchey i neorganicheskoy khimii AN UkrSSR.
(Silver--Analysis) (Drinking water)

KUL'SKIY, L.A.; NIKITINA, S.V.; SLIPCHENKO, V.A.

Preserving drinking water by means of silver electrolytic solutions. Ukr.khim.zhur. 28 no.8:981-986 '62. (MIRA 15:11)

1. Institut obshchey i neorganicheskoy khimii AN UkrSSR.
(Drinking water)
(Silver)

L 12811-63 EWP(q)/EWT(m)/BDS AFFTC/ASD JD/JG

ACCESSION NR: AP3000780 S/0070/63/008/003/0453/0454

57

56

AUTHOR: Kuz'min, R. N.; Nikitina, S. V.

21

TITLE: Structure of compounds of rare-earth metals with antimony and bismuth having the composition AB

21 21

SOURCE: Kristallografiya, v. 8, no. 3, 1963, 453-454

TOPIC TAGS: semiconductor, superconductivity, rare earths, Sb, Bi, NaCl, covalent bond

ABSTRACT: This study was made to discover new compounds of these elements. The results add to the information of crystal chemistry of superconductivity in binary alloys. The compounds obtained represent NaCl structure, as indicated by a comparison of calculated and measured intensities on x-ray photographs. The interatomic distances suggest the presence of a mixed ionic-covalent chemical bond. In view of this and of the results in general, the authors conclude that superconductivity should not appear in compounds of rare earths with Sb and Bi in which the compositional formula is AB. The presence of covalent bonds suggests that the compounds are semiconductors, and the preservation of NaCl structure makes them important subjects of investigation relative to their electrical and magnetic properties. Orig. art. has: 1 table.

Card 1/1 Moscow St. Univ

KUL'SKIY, L.A.; SOTNIKOVA, Ye.V.; NIKITINA, S.V.; SLIPCHENKO, V.A.

Long-term storage of drinking water containing silver ions.
Gig. i san. 28 no.1:99-102 Ja'63. (MIRA 16:7)

1. Iz Instituta obshchey i neorganicheskoy khimii AN SSSR.
(WATER—PURIFICATION) (SILVER IONS)

GENEL', S. V.; KOVAL', P. M.; NIKITINA, T. A.

Looms

Pasting a fibrous cover on shuttles Tekst. prom. no. 5, 1952

*Monthly List of Russian Accessions*, Library of Congress, August, 1952. UNCLASSIFIED.

NIKITINA, T.A.

The ATM-160-5 automatic loom for manufacturing terry cloth. Biul.-tekh.-ekon.inform.Gos.nauch.-issl.inst.nauch. i tekh.inform. no.6:55-56 '62. (MIRA 15:7)

(Looms)

BOYARCHIKOV, A.I.; NIKITINA, T.A.; TOPILINA, V.S.

Increasing the speed of the main shaft of AT-100-5M looms. Tekst.prom. 23 no.11:59-60 N '63. (MIRA 17:1)

1. Starshiy inzh. Moskovskogo soveta narodnogo khozyaystva (for Boyarchikov). 2. Vedushchiy inzh. Vsesoyuznogo nauchno-issledovatel'skogo instituta legkogo i tekstil'nogo mashinostroyeniya (for Nikitina). 3. Starshiy tekhnik Vsesoyuznogo nauchno-issledovatel'skogo instituta legkogo i tekstil'nogo mashinostroyeniya (for Topilina).

BOYARCHIKOV, A.I.; NIKITINA, T.A.; TOPILINA, V.S.

Increasing the speed of the main shaft of AT-100-5M looms. Tekst.prom. 23 no.11:59-60 N '63. (MIRA 17:1)

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NIKITINA, T.A.

Results of the organization of control of dermatomycoses in rural medical stations. Vest. vener., Moskva no. 4:43-44 July-Aug 1952.
(CLML 23:3)

1. Of the Mycological Laboratory (Head -- T. A. Nikitina), Kiev Scientific-Research Dermato-Venereological Institute (Director -- Prof. A. P. Lavrov).

NIKITINA, T. A.

Skin - Diseases

Cases of dermatomycosis in Kiev in post-war years. Vest. Ven. i Derm. [illegible]

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**NIKITINA, T.A.**

Clinical aspect and mycology of microsporia caused by Microsporum ferrugineum. Vest.ven.i derm. no.6:12-16 N-D '53. (MLRA 6:12)

1. Iz mikologicheskoy laboratorii (zaveduyushchiy T.A.Nikitina) Kiyevskogo nauchno-issledovatel'skogo dermato-venerologicheskogo instituta (direktor - professor A.P.Lavrov).
(Medical mycology) (Skin--Diseases)

NIKITINA, T.A.

MEL'NIK, M.I.; NIKITINA, T.A., kandidat meditsinskikh nauk; BABASEVA, Ye.P.; FOKINA, A.I.; KONONOV, O.K.; SEVERIN, A.V.

Treatment of mycoses of the scalp with Lesovykh solutions No.1 and No.2 without using X rays. Vest. ven. i derm. no.5:21-22 S-O '54. (MLRA 7:11)

1. Iz Kiyevskogo dermato-venerologicheskogo instituta (dir. G.Ye. Koryakin) i Kiyevskogo gorodskogo dispansera (glavnyy vrach A.S. Ivanov)

(HEAD, diseases,
fungus dis., chemother.)
(FUNGUS DISEASES,
scalp, chemother.)

NIKITINA, T.A., kandidat meditsinskikh nauk

[redacted]

Fungous flora of some Ukrainian Provinces west of the Dnieper during the postwar period. Vest.ven. i derm. no.2:54-55 Mr-Ap '55.
(MLRA 8:5)

1. Iz mikrologicheskoy laboratorii Kiyevskogo venerologicheskogo instituta.
(UKRAINE -- FUNGI, PATHOGENIC)

NIKITINA, T. A., kend. med. nauk.

Human mycosis. Vest. ven. i derm. no.5:50-52 S-O '55 (MLRA 9:1)

1. Iz milologicheskoy laboratorii (zav.-kandidat meditsinskikh nauk T. A. Nikitina)Kiyevskogo nauchno-issledovatel'skogo kozhno-venerologicheskogo instituta (dir. G. Ye.Koryakin, nauchnyy rukovoditel'-prof. M. M. Kuznets)
(FUNGUS DISEASES
pathol, diag. & ther.)

NEKACHALOV, V.Ya.; NIKITINA, T.A.

Analysis of some cases of increased resistance to antibiotics developing in medical workers. Eksp. i klin. issl. po antibiot. 1:376-382 '58. (MIRA 15:5)
(ANTIBIOTICS) (ALLERGY)

NIKITINA, T.A.; FROLOVA, M.A.

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(ANTIBIOTICS) (ALLERGY)

NEKACHALOV, V.Ya.; MARGOLIN, A.M.; NIKITINA, T.A.; LISOVSKAYA, N.D.;
KHARENKO, V.I.; MAL'GINA, V.G.

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(MONILIASIS) (ANTIBIOTICS--TOXICOLOGY)

**NIKITINA, T.A.**, kand.med.nauk

Pharyngomycosis. Vest. otorin. 22 no.1:46-48 Ja-F '60.
(MIRA 14:5)

1. Iz mikologicheskoy kliniki (zav. - dotsent V.Ya.Nekachalov) Leningradskogo instituta antibiotikov (nauchnyy rukovoditel' kliniki - prof. A.N.Araviyskiy).
(PHARYNX—DISEASES)

**NIKITINA, T.A.**

Anaphylactic reaction fo the intradermal adminstration of adrenocorticotropic hormone. Vest.derm.i ven. 34 no.3:12-14 My-Je '60. (MIRA 13:10)

(ACTH) (ANAPHYLAXIS)